

# *International Review of Mechanical Engineering (IREME)*

## Contents

<b>Analytical Energetic Approach for Predictive Generation of Dynamic Biped Walking - Use of Average Energies</b> <i>by S. Alfayad, O. Bruneau, F. B. Ouezdou, Moustafa A. Fouz, P. F. Doubliez</i>	461
<b>Compressible Subsonic Flow in Gas Turbine Annular Diffusers</b> <i>by Meiriele Abreu Aharenga, Cláudia R. Andrade, Edson L. Zaporoli</i>	474
<b>Dynamic Modeling and PID Control of an Underwater Robot Based on the Hardware-in-the-Loop Method</b> <i>by Ruben D. Hernandez B., Pablo A. Mora, Oscar F. Avilés, Janito V. Ferreira</i>	482
<b>The Influence of Thermal Grashof Number, Reynolds Number and Schmidt Number on the Behaviour of a Flow in a Porous Medium</b> <i>by H. Oudrhiri, K. Gueraoui, A. Elbouzidi, M. Sammouda, Y. Belkassmi</i>	491
<b>Statistical Approach of Bearing Degradation</b> <i>by M. Bourdim, A. Bourdim, S. Kerrouz, A. Benamar</i>	496
<b>Influence of Critical Process Parameters on the Quality of Friction Stir Welded Nylon 6</b> <i>by M. S. Youssif, M. A. El-Sayed, A. M. Khourshid</i>	501
<b>An Analytical Solution for Two-Layered Composite Beams with Imperfect Shear Interaction</b> <i>by Ákos József Lengyel, István Ecsedi</i>	508
<b>Revisiting the Von Kármán Problem with an Analytical Solution</b> <i>by Sebastião Cardoso, José M. Balthazar, Ezio C. Garcia</i>	518
<b>Severity of the Residual Stress Depending on the Width of Welding in the Rails</b> <i>by O. Bouazaoui, A. Chouaf</i>	523
<b>Effect of Organics Corrosion Inhibitors on the Corrosion of 304SS in 3.5% NaCl</b> <i>by F. Gapsari, R. Soenoko, A. Suprapto, W. Suprapto</i>	531
<b>Development of Chitosan Coating: Study of the Effect of Roughness on the Mechanical and Morphological Performances</b> <i>by F. Z. Bougueraa, S. Ettaqi</i>	537
<b>Jerk Trajectory Planning for Assistive and Rehabilitative Mechatronic Devices</b> <i>by F. Aggeleri, A. Borboni, N. Pellegrini</i>	543
<b>Performances and Emissions Characteristics of Three Main Types Composition of Gasoline-Ethanol Blended in Spark Ignition Engines</b> <i>by Marith E. N. Paloboran, I. Nyoman Sutantra, Bambang Sudarmanta</i>	552

# Effect of Organics Corrosion Inhibitors on the Corrosion of 304SS in 3.5% NaCl

F. Gapsari, R. Soenoko, A. Suprapto, W. Suprapto

**Abstract** – This work is a preliminary study with the purpose to find some information about the effectiveness of Ceraflava (CF) extract as inhibitor in 3.5 %NaCl, CF or yellow bees wax from honey farm used as corrosion inhibitor. In particular the purpose of this study is to determine the ability of CF in inhibiting the corrosion rate. The inhibition of the CF extract on304SS in 3.5% NaCl is carried out with potentiodynamic polarization, EIS and SEM. The findings show that CF can be used as corrosion inhibitor in SS 304 even though CF can only be categorized as moderate performance of inhibitors. However, it was observed that inhibitors can be adsorbed on the 304SS surface so it inhibits the corrosion rate. Based on the potentiodynamic polarization and EIS results, the inhibition efficiencies are 59.77% and 53.84% in 0.2g/L inhibitor concentration. Potentiodynamic polarization curves indicate the mixed type of the inhibitors. The adsorption of the inhibitors on the 304SS surface was found to obey to Langmuir adsorption isotherm. Thermodynamic adsorption parameter were calculated and discussed. It is expected this research may be recommended in the use of CFCs in saline solution or other solutions on the science of corrosion controlling using organic inhibitors. Copyright © 2016 Praise Worthy Prize S.r.l. - All rights reserved.

**Keywords:** 304SS, Electrochemical Technique, Inhibitors, NaCl Solution, SEM

## I. Introduction

Type 304SS is the most versatile and widely used of all the stainless steels [1]-[43]. The 304SS has a good corrosion resistance [1]. The 304SS is the type of stainless steel mainly used for canned food, kitchen utensils, medical devices and automotive industry [2]. However, the disadvantage of this type is that it stimulates local corrosion when exposed to chloride or located in the environment containing chloride ( $\text{Cl}^-$ ) ion [3]. The environments with the most frequent contacts with metal are water, environment containing the  $\text{Cl}^-$  ion and acid environment. Chloride ion is the most aggressive ion to cause metal corrosion especially on the 304SS. However, the resistance to pitting corrosion of 304SS in solutions containing  $\text{Cl}^-$  is not good enough and it adversely influences both the service life and integrity of structures made by this material [1]. Some studies on the 304SS corrosion have been widely conducted [1], [4], [5]. Similarly, many efforts have been taken in order to increase the 304SS resistance towards NaCl [6]-[9]. The inhibitor is the most effective method in preventing the corrosion of metal surfaces by aggressive solutions. Organic compounds act as a good inhibitors due to their heteroatom structures such as sulfur, nitrogen and oxygen [10]–[14]. Various organic compounds have been studied as corrosion inhibitors for 304SS in different solution [15]-[17].

In the present study, Ceraflava (CF) or yellow bees wax by honey farm is used as corrosion inhibitor.

CF is generally used as a remedy for dry lips or skins. It contains flavanoid and antioxidant. The aim of this paper is to study the effectiveness of the inhibitor addition on the 304SS in 3.5%NaCl solution.

## II. Materials and Methods

### II.1. Cera Flava Extraction

After the honey and the honeycomb were separated, the remaining is CeraFlava.200 grams of CF were mashed and extracted with 200 ml of 99% ethanol (Merck). The mixture was filtered and then refluxed for 4 hours at 50°C.

### II.2. FTIR Testing

Shimadzu IR Prestige-21 type of Fourier transform infrared (FTIR) spectrometer instrument was used with the following resulting characteristics: IR-spectral in range 4000-400 ( $\text{cm}^{-1}$ ) with 4  $\text{cm}^{-1}$ resolution and KBr matrix with a scan rate of 32 scans per minute.

The functional groups in CF extract are determined by FTIR.

### II.3. Materials

The chemical composition (wt. %) of 304SS is 0.04%C, 0.92 % Mn, 0.52 % Si, 0.030% P, 0.002 % S, 18.15 % Cr, 9.58% Ni and Bal.Fe.

The density of 304SS was 7.9 g/cm<sup>3</sup>. The 304SS was added to epoxy resin with a geometric exposed surface area measuring 1 cm<sup>2</sup> and connected to the electrolyte.

Before the experiments, the specimens were polished with 200, 400, 600, 1000, 1500, and 2000 grades of SiC paper.

#### II.4. Electrochemical Measurement

Corrosion testing was carried out using the G-31 American Standard and Testing (ASTM) [18].

The variable used in the study is an independent one in the form of inhibitor concentration, namely 0, 0.1, 0.2, 0.3, and 0.4 g/L of CF extract.

The room temperature was 25°C. The electrochemical testing was carried out with the PGSTAT 128N AUTOLAB device. The prepared specimen was made into an electrochemical cell with platinum as the counter electrode (CE), Ag/AgCl (KCl 3 M) as the reference electrode (RE), and 304SS specimen as the working electrode (WE). The three electrodes were immersed in batch for 3 hours. The measurement of polarization in -2 V until 1.5 V range corrosion potential (OCP) with scan rate has given a value of 20 mV/s. The inhibition efficiency was measured with the following Eq. (1) [19]:

$$EI(\%) = \frac{I_{cor} - I_{cor(i)}}{I_{cor}} \times 100\% \quad (1)$$

where  $I_{cor}$  and  $I_{cor(i)}$  indicate the corrosion current density with and without addition of inhibitors.

The *Electrochemical Impedance Spectroscopy* (EIS) measurement was carried out at the frequency of 100 kHz to 10 MHz and 5 mV as amplitude using AC signal at Open Circuit Potential (OCP). The inhibition efficiency with the EIS method was formulated with the following Eq. (2) [29]:

$$EI(\%) = \frac{R_{ct(i)} - R_{ct}}{R_{ct(i)}} \times 100 \quad (2)$$

### III. Result and Discussion

#### III.1. FTIR Analysis

Based on the FTIR testing, the function of the CF extract can be found. Fig. 1 shows the function of the CF extract (see Fig. 1). The result of FTIR analysis is shown in Fig. 1. There are some absorbed spectrum with high intensity at some wave numbers, namely 3369.41 cm<sup>-1</sup> that is detected as the absorption of phenol functional group. The absorption of CH alkane appears at 2925.81 cm<sup>-1</sup>. The CC alkene is shown at 1618.17 cm<sup>-1</sup>.

The other absorption are shown at 1454.23 cm<sup>-1</sup> of the CH alkane functional group, and 1049.20 cm<sup>-1</sup> as the CO alcohol functional group. It shows that the CF extract has C – H, C = O, C = C, and O – H as its functional groups.

The functional groups match phenol and aromatic compounds.

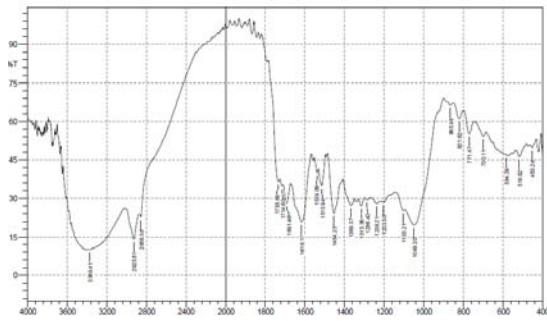


Fig. 1. Result of FTIR Testing of CF Extract

#### III.2. Potentiodynamic Polarization Studies

Polarization testing was carried out towards CF extract concentration variation. The results of inhibition efficiency measurement using the polarization method are presented in Table I.

Polarization testing was carried out towards CF extract concentration variation. The results of inhibition efficiency measurement using the polarization method are presented in Table I. Fig. 2 shows tafel plots with various concentration. The optimum inhibition efficiency obtained was 59.77% by addition of 0.2 g/L inhibitor.

TABLE I  
INHIBITION EFFICIENCY MEASUREMENT WITH  
THE POLARIZATION METHOD

CF extract (g/L)	$\beta_a$ (V/dec)	$B_c$ (V/dec)	$E_{corr}$ (V)	$I_{corr}$ (A/cm <sup>2</sup> )	IE (%)
0	0.156	5.440	-0.341	$2.61 \times 10^{-6}$	0
0.1	0.109	0.149	-0.378	$1.19 \times 10^{-6}$	54.40
0.2	0.140	0.194	-0.370	$1.05 \times 10^{-6}$	59.77
0.3	0.085	0.180	-0.198	$1.53 \times 10^{-6}$	41.37
0.4	0.124	0.302	-0.344	$1.62 \times 10^{-6}$	37.93

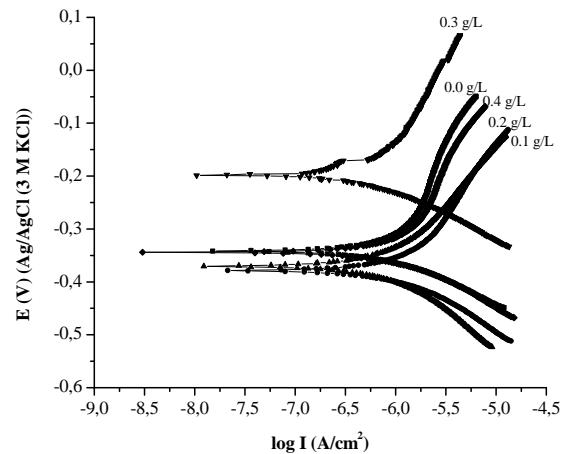


Fig. 2. Tafel plots with various concentration of CF extract in 3.5% NaCl

The inhibition efficiency increased at 0.1 to 0.2 g/L addition of inhibitors. Then, the efficiency decreased at 0.3 g/L addition of inhibitors. The addition of inhibitors as much as 0.1; 0.2 g/L, the corrosion potential ( $E_{corr}$ ) show a tendency to anodic (Fig. 2). It shows that the CF extract controls the anodic reaction by forming complex molecules in the anodic area on the 304SS surface [20].

On the other hand, change into cathodic occurred with the addition of 0.3 g/L. The cause of the change was the formation of complex molecules in the cathodic area of the 304SS surface [20]. The Ecorr values of the solution with inhibitors increased and decreased randomly compared to a solution without inhibitors. It indicates that CF extract is categorized as mixed type inhibitors [21], [22]. The optimum inhibition efficiency occurred in 0.2 g/L and decreased with the addition of inhibitor concentration; a possible cause of the tendency can be physisorption. The more addition of CF extract caused a saturated solution and released the protective film so that the efficiency decreased. CF extract which was added and functioned as the inhibitors and had optimum concentration, when added beyond its optimum concentration, will saturate and may cause formed layer separated. This was caused by the interaction among inhibitor molecules larger than inhibitor interaction on the metal surface. The inhibition efficiency decreased at higher concentration caused by desorption molecules from the metal surfaces [23], [24]. The saturated solution initiated interaction between metal and surface stopped and caused the metal surface not to be covered therefore decreasing the inhibition efficiency.

The change of anodic ( $\beta_a$ ) and cathodic ( $\beta_c$ ) tafel constants is irregular (Table I). The irregularity of  $\beta_c$  shows that the inhibitor is not optimum to decrease the cathodic reaction in the metal [25]. On the other hand, the irregularity of  $\beta_a$  value indicates that there is physisorption on the metal surface.

The CF extract has a moderate inhibition capacity in all the variations of concentration. The inhibitor capacity to decrease the corrosion is different from another depending upon its ability to form complex molecules. Upon the relative solubility, the result can either inhibit or catalyze the further dissolution of metal [26]. The inhibition is definite and depends on the characteristics of metal and corrosive environment [27]. Meanwhile, the corrosion inhibitors are opted for inhibited solubility of the fluid [28]. The immediate absorption from the functional group forces the molecules of CF extract to move horizontally using the surface of orientated metal.

The immediate absorption causes each functional group interact on its own.

The corrosion rate of 304SS NaCl solution is set off through two main procedures, namely formation and build-up of an iron oxide layer by pitting destruction [29]. Due to the corrosion mechanism of stainless steel in a saline mode, at the anode regions, stainless steel has the following potential dissolution reaction:  $\text{Fe} \rightarrow \text{Fe}^{2+} + 2\text{e}^-$ , occurring adjacently to the surface of the cathode reaction:  $\frac{1}{2}\text{O}_2 + \text{H}_2\text{O} + 2\text{e}^- \rightarrow 2\text{OH}^-$  [8]. The missing inhibitor brings about an active behavior, without the formation of a passive layer showing offensive formation of corrosion product layer. When inhibitor is added, there is a tendency that the Ecorr values move toward more active value while the Icorr values decrease. The anodic shows an active behavior, similar to what is observed for non-inhibitor solution.

The behavior might be caused by the absence of inhibitor that dissolves in water [30]. The inhibitor may not be able to be put directly into the corrosive solution.

In cathodic, the current density is limited due to the oxygen reduction. It can be seen that the addition of inhibitor has no significant effect in the cathodic at polarization curve because the limitation of current density. The change to the cathodic occurs only at the addition of 0.3 g/L as the inhibitor after surpassing the optimum concentration to increase inhibition efficiency.

There is a possibility that it causes the maximum inhibition capacity at 59.77%.

### III.3. Electrochemical Impedance Spectroscopy (EIS)

The impedance diagrams represented in Nyquist plot obtained at OCP are presented in Fig. 3. The parameters of the EIS testing are described in Table II. They are the transfer resistance of  $R_{ct}$ , the solution resistance ( $R_s$ ) and the Constant Phase Element(CPE). The interaction between the surface of metal and the solution causes a transfer charge between them measured as  $R_{ct}$ .

The corrosion behavior of 304SS frequency in different variations of solution concentration is shown by the Nyquist plot as seen in Fig. 3. The partial circle shows the characteristics of solid electrode and it has irregular and non-homogenous electrode [31], [32].

The Nyquist plots of 304SS at the various concentration of inhibitor exhibit similar patterns. This occurrence specifies the inhibition mechanism (a charge-transfer process between the inhibitor and the metal surface) and it is also alike [15]. In this case, the result of the EIS test shows similar result to the potentiodynamic polarization with the highest inhibition efficiency at the concentration of 0.2 g/L (Section III.1).

TABLE II  
IMPEDANCE PARAMETER OF THE 304SS IN 3.5% NaCl

CF extract (g/L)	$R_{ct}$ ( $\Omega$ )	$R_s$ ( $\Omega$ )	CPE ( $\mu\text{F}$ )	$\chi^2$	IE (%)
0.0	39.924	1.012	154.410	0.036907	0
0.1	72.868	1.212	68.437	0.011062	45.21
0.2	86.487	1.877	57.212	0.037338	53.84
0.3	62.178	1.779	68.166	0.009212	35.79
0.4	52.691	1.471	65.258	0.015423	24.23

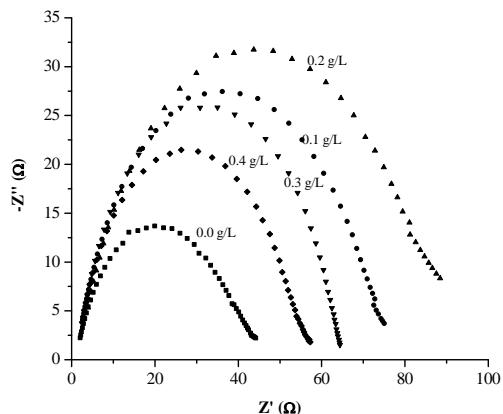


Fig. 3. The Nyquist plot of the 304SS in 3.5%NaCl with the cf extract concentration variation

Fig. 4 shows the corresponding circuit diagram used to fit the EIS in which the calculation of the impedance parameters was simulated using the corresponding circuit. The interaction between the metal and the solution was illustrated.

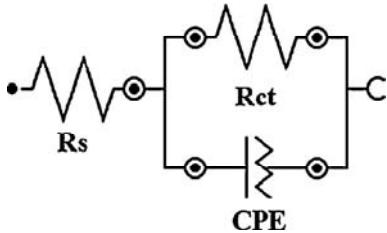


Fig. 4. Equivalent circuit used to fit the EIS diagrams

#### III.4. Adsorption Isotherm

The mechanism of the corrosion protection may be explained on the basis of adsorption behavior [33].

The interaction information between the inhibitor molecule and metal surface can be provided by the adsorption isotherm [34]. Generally, organic molecules can prevent the corrosion on metal using the adsorption. The type of adsorption depends on chemical molecule composition, temperature and electrochemical potential of the metal or interface between solution and metal. The measurement of isothermal adsorption of the findings is adjusted to match various isothermal equations, namely Frumkin, Langmuir, Temkin, Freundlich, Bockris-Swinkels and Flory-Huggins [35], [36]. The best fitted result of the study follows the Langmuir isothermal adsorption (equation (3)) shown in Fig. 5:

*Langmuir's Equation:*

$$\frac{C}{\theta} = \frac{1}{K_{ads}} + C \quad (3)$$

where  $\theta = \frac{EI}{100}$ ,  $\theta$  = surface coverage and  $K_{ads}$  = adsorption equilibrium constants.

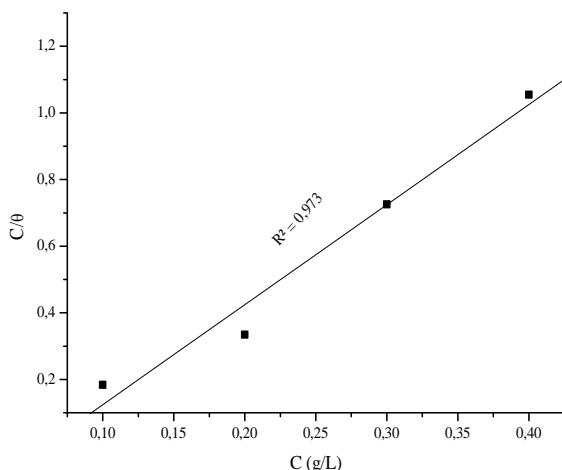


Fig. 5. Langmuir adsorption isotherm

The free energy of adsorption ( $\Delta G_{ads}^\circ$ ) was calculated by the equilibrium constant of adsorption ( $K_{ads}$ ) using Eq. (4):

$$\Delta G_{ads}^\circ = -2.303RT \log(K_{ads} \times 1000) \quad (4)$$

where 1000 is the concentration of water in the solution (g/L), R is the universal gas constant and T is the absolute temperature.

The values of free energy are negative as well as less than - 40 kJmol<sup>-1</sup>, indicating the spontaneous physical adsorption of the inhibitors on the metal surface [37].

It gives the same conclusion as the findings of polarization and EIS. The negative values of  $G_{ads}$  showed that the adsorption of inhibitor molecules on the surface occur spontaneously [38]. The calculation of the enthalpy value ( $\Delta H_{ads}^\circ$ ) is the Van't Hoff's formula in Eq. (5):

$$\ln K_{ads} = \frac{-\Delta H_{ads}^\circ}{RT} \quad (5)$$

So that the entropy adsorption ( $\Delta S_{ads}^\circ$ ) can be measured using Eq. (6):

$$\Delta G_{ads} = \Delta H_{ads}^\circ - T\Delta S_{ads}^\circ \quad (6)$$

Therefore, the thermodynamic parameter of inhibitor adsorption can be seen in Table III.

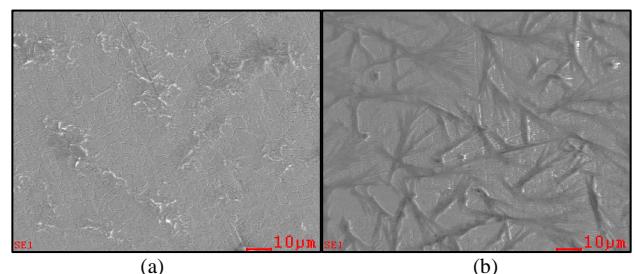
The positive value of  $\Delta H_{ads}^\circ$  shows that the process of dissolution is exothermic, while, the positive value of  $\Delta S_{ads}^\circ$  shows that the adsorption process followed with an increase in the entropy, which is the driving force for the adsorption inhibitor to the metal surface [39].

#### III.5. SEM Analysis

The results of SEM are illustrated in Figs. 6 showing micro photograph of the specimen without inhibitor and specimen with the CF extract inhibitor with 0.2 g/L concentration. Figs. 6(a) and 6(b) show the images of the specimen without and with inhibitor, respectively.

On Fig. 6(a), it can be seen a small hole on the 304SS surface. Meanwhile, the solution with inhibitor is smoother even though there are still small holes that indicate corrosion.

Therefore, there should be a synergic effect effort to improve the performance of inhibitor.



Figs. 6. Scanning Electron Microscopic images of 304SS  
(a) absence inhibitor (b) presence inhibitor in 3.5% NaCl

## IV. Conclusion

Based on the results, the CF extract as corrosion inhibitors works by physisorption. The inhibition capacity of the CF extract is relatively low.

The highest inhibition efficiency occurs in the addition of 0.2 gr/L of CF extract as much as 59.77% and 53.84% with the potentiodynamic polarization and EIS measurement, respectively.

The electrochemical testing showed that the CF extract can carry out adsorption on the 304SS surface, but it is unable to form a protective layer either simultaneously or thoroughly on the surface. There should be more efforts for synergistic effects so that the performances of inhibitor will increase.

## References

- [1] C. Pan, L. Liu.,Y. Liu,F. Wang, Pitting corrosion of 304SS nanocrystalline thin film.*Corrosion Science*, Vol. 73, pp. 32-43, 2013.  
<http://doi:10.1016/j.corsci.2013.03.022>
- [2] P. Fellows, B.Axtell, *Appropriate Food Packaging; Material and Methods for Small Business*, Essex U.K: ITDG Publishing.
- [3] J. Baszkiewicz, M. Kamiński, A. Padgrsky, J. Jagieki, G. Gawlik, Pitting corrosion resistance of silicon-implanted stainless steels, *Corrosion Science*, Vol. 33, pp. 815-818, 1992.  
[http://doi:10.1016/0010-938X\(92\)90114-I](http://doi:10.1016/0010-938X(92)90114-I)
- [4] Y. Wang,W. Wang, Y.Liu, L. Zhong,J. Wang,Study of localized corrosion of 304 stainless steel under chloride solution droplets using the wire beam electrode, *Corrosion Science*, Vol. 53, pp. 2963-2968.  
<http://10.1016/j.corsci.2011.05.051>
- [5] C. Pan, L. Liu, Y. Li, S.H. Wang, F.H. Wang, Passive film growth mechanism of nanocrystalline 304 stainless steel prepared by magnetron sputtering and deep rolling techniques, *Electrochimica Acta*, 56, 7740-7748, 2011.  
<http://doi:10.1016/j.electacta.2011.05.106>
- [6] Y. Zhang, X. Yin,F. Yan, Effect of halide concentration on tribocorrosion behaviour of 304SS in artificial seawater, *Corrosion Science*, Vol. 99, pp. 272-280, 2015.  
<http://doi:10.1016/j.corsci.2015.07.017>
- [7] F. Yu,S. Chen, Y.Chen, H. Li, L.Yang, Y.Chen, Y. Yin,Experimental and theoretical analysis of polymerization reaction process on the polydopamine membranes and its corrosion protection properties for 304 Stainless Steel. *Journal of MolecularStructure*, Vol. 982, pp. 152-160, 2010.  
<http://doi:10.1016/j.molstruc.2010.08.021>
- [8] F. Yu, S.Chen, H.Li, L.Yang, Y.Yin,Application of Self Assembled 6-aminohexanol layers for corrosion protection of 304 stainless steel surface, *Thin Solid Films*, Vol. 520, pp. 4990-4995, 2012.  
<http://doi:10.1016/j.tsf.2012.03.006>
- [9] S. Y. Arman, H. Omidvar, S.H.Tabaian,M. Sajjadnejad, Sh.Fouladvand, Sh.Afshar,Evaluation of nanostructured S-doped TiO<sub>2</sub> thin films and their photoelectrochemical application as photoanode for corrosion protection of 304 stainless steel,*Surface and Coating Technology*, Vol.251, pp. 162-169, 2014.  
<http://doi:10.1016/j.surcoat.2014.04.020>
- [10] R.O. Ramos, A.Battistin, R.S. Gonçalves,Alcoholic *Mentha* extracts as inhibitors of low-carbon steel corrosion in aqueous medium. *Journal Solid State Electrochemistry*. 16, 747-752, 2012.  
<http://DOI 10.1007/s10008-011-1422-8>
- [11] C. Wang, S. Chen, S. Zhao, Inhibition effect of ac-treated, mixed self-assembled film of phenylthiourea and dodecanethiol on copper corrosion, *Journal of Electrochemistry Society*, Vol.151, pp. B11-B15, 2004.  
<http://doi:10.1149/1.1630595>
- [12] G.K. Gomma, M. H.Wahdan,Effect of temperature on the acidic dissolution of copper in the presence of amino acids; *Material Chemistry and Physics*,Vol. 39, pp. 142-148, 1994.  
[http:// doi:10.1016/0254-0584\(94\)90191-0](http:// doi:10.1016/0254-0584(94)90191-0)
- [13] M. Kendig,S. JeanjaquetCr(vi) and ce(iii) inhibition of oxygen reduction on copper; *Journal of Electrochemistry Society.*, 149, B47-B51, 2004.  
<http://doi:10.1149/1.1430717>
- [14] K.F. Khaled, N. Hackerman,*Ortho*-substituted anilines to inhibit copper corrosion in aerated 0.5 M hydrochloric acid, *Electrochimica Acta*, Vol. 49, pp. 485-489, 2004.  
<http://doi:10.1016/j.electacta.2003.09.005>
- [15] F. Kurniawan, K.A.Madurani, Electrochemical and optical microscopy study of red pepper seed oil corrosion inhibition by self-assembled monolayers (SAM) on 304SS, *Progress on Organic Coating*, Vol. 88, pp. 256-262, 2015.  
<http://dx.doi.org/10.1016/j.porgcoat.2015.07.010>
- [16] F. Gapsari, R.Soenoko, A.Suprapto, Bee wax propolis extract as eco-friendly corrosion inhibitors for 304ss in sulfuric acid, *International Journal of Corrosion*, 2015.  
<http://dx.doi.org/10.1155/2015/567202>
- [17] M. Urgen, A.F.Cakir,The Effect of Molybdate ions on the temperature dependent pitting potential of austenitic Stainless steel in neutral chloride solutions, *Corrosion Science*, Vol.32, pp. 835-852, 1991.  
[http:// doi:10.1016/0010-938X\(91\)90028-N](http:// doi:10.1016/0010-938X(91)90028-N)
- [18] ASTM G1-72; Practice for Preparing, Cleaning and Evaluating Corrosion Test Specimens; ASTM, 1990.
- [19] M.I. Awad,Eco friendly corrosion inhibitors: inhibitive action of quinine for corrosion of low carbon steel in 1MHCl, *Journal of Applied Electrochemistry*, Vol. 36, pp. 1163-1168, 2006.  
<http://dx.doi.org/10.1007/s10800-006-9204-1>
- [20] S. Rajendran, M.Agasta, R.B. Devi, B.S. Devi,K. Rajam, J. Jeyasundari,Corrosion inhibition by an aqueous extract of Henna leaves (*Lawsonia inermis* L),*ZastitaMaterijala*, Vol.50,pp. 77-84, 2009.
- [21] K.P.V.Kumar, M.S.N Pillai, G.R. Thusnavis, Pericarp of the Fruit of *Garcinia Mangostana* as Corrosion Inhibitor for Mild Steel in Hydrochloric Acid Medium.*Portugaliae Electrochimica Acta*, Vol. 28, pp. 373-383, 2010.  
<http:// DOI: 10.4152/pea.201006373>
- [22] A.S. Fouda, H.A.Mostafa, H.M. El-Abbasy, Antibacterial drugs as inhibitors for the corrosion of stainless steel type 304 in HCl solution, *Journal of Applied Electrochemistry*, Vol. 40, pp. 163-173, 2010.  
<http://DOI: 10.1007/s10800-009-9992-1>
- [23] R. Solmaz,Investigation of the inhibition effect of 5-((E)-4-phenylbuta-1,3 dienylideneamino)-1,3,4-thiadiazole-2-thiol Schiff base on mild steel corrosion in hydrochloric acid, *Corrosion Science*, Vol.52, pp. 3321-3330, 2010.  
<http://doi:10.1016/j.corsci.2010.06.001>
- [24] R. Solmaz,Investigation of adsorption and corrosion inhibition of mild steel in hydrochloric acid solution by 5-(4-Dimethylaminobenzylidene) rhodanine, *CorrosionScience*, Vol. 79, pp. 169-176, 2014.  
<http://doi:10.1016/j.corsci.2013.11.001>
- [25] A.K. Mohamed, H.A.Mostafa, G.Y. El-Awady, A.S.Fouda,Use of 1- phenylamino 3-(4-[henythiosemi carbazole]-butane-i3-dione derivatives as corrosion inhibitors for c-steel in acidic chloride solutions, *Portugaliae Electrochimica Acta*, Vol. 18, pp. 99-112, 2000.
- [26] D.M. Ortega-Toledo,J.G. Gonzalez-Rodriguez, M. Casales, A.Caceres, L.Martinez, Hydrodynamic effects on the co<sub>2</sub> corrosion inhibition of x-120 pipeline steel by carboxyethyl-imidazoline, *International Journal of Electrochemical Science*, Vol. 6, pp. 778-792, 2011.
- [27] R.T. Loto, C.A. Loto, Effect of p-phenylenediamine on the corrosion of austenitic stainless steel type 304 in hydrochloric acid, *International Journal of Electrochemical Science*, Vol. 7, pp. 9423-9440, 2012.
- [28] Rahimi, Inorganic and Organometallic Polymers: A Review, *Iranian Polymer Journal*, Vol. 13, pp. 149-164.
- [29] L.Cáceres, T. Vargas, L. Herrera, Influence of pitting and iron oxide formation during corrosion of carbon steel in unbuffered NaCl solutions, *Corrosion Science*, Vol. 51, pp. 971-978, 2009.

- http://doi:10.1016/j.corsci.2009.02.021
- [30] L.M. Rivera-Grau, M.Casales, I.Regla, D.M. Ortega-Toledo, J.A.Ascencio-Gutierrez, J. Porcayo-Calderon, L. Martinez-Gomez, Effect of organic corrosion inhibitors on the corrosion performance of 1018 carbon steel in 3% NaCl solution, *International Journal of Electrochemical Science*, Vol.8, pp. 2491-2503, 2013.
- [31] K. Juttner,Electrochemical impedance spectroscopy (EIS) of corrosion processes on inhomogeneous surfaces, *Electrochimica Acta*, Vol. 35, pp. 1501-1508, 1990  
http://doi:10.1016/0013-4686(90)80004-8
- [32] M. H. Hussin, M. J. Kassim,The corrosion inhibition and adsorption behavior of *Uncaria gambir* extract on mild steel in 1 M HCl; *Material Chemistry and Physics*, Vol. 125, pp. 461-468, 2010.  
http://doi:10.1016/j.matchemphys.2010.10.032
- [33] R.F.V. Villamil, P. Corio, J.C. Rubin, S.M.I. Agostinho,Effect of sodium dodecylsulfate on copper corrosion in sulfuric acid media in the absence and presence of benzotriazole, *Journal of Electroanalytical Chemistry*, Vol. 472, pp. 112-119, 1999.  
http://doi:10.1016/S0022-0728(99)00267-3
- [34] I.B. Obot, N.O. Obi-Egbedi, S.A.Umoren,Experimental and theoretical investigation of clorotrimazole as corrosion inhibitor for aluminium hydrochloric and effect of iodide ion addition, *De Pharma Chemica*,Vol. I, pp. 151-161, 2009.
- [35] A.K. Maaya, N. Al-Rawashdeh, Inhibition of acidic corrosion of pure aluminum by some organic compounds, *Corrosion Science*, 46, 1129-1140, 2004.  
http://doi:10.1016/j.corsci.2003.09.009
- [36] X.Li, S.Deng, H. Fu, T. Li, Adsorption and inhibition effect of 6-benzylaminopurine on cold rolled steel in 1.0 HCl, *Electrochimica Acta*, Vol. 54, pp. 4089-4098 , 2009.  
http://doi:10.1016/j.electacta.2009.02.084
- [37] A. Nahle, I. Abu-Abdoun, I. Abdel-Rahman, M.Al-Khayat, UAE neem extract as a corrosion inhibitor for carbon steel in HCl Solution, *International Journal of Corrosion*, 2015.  
http://dx.doi.org/10.1155/2010/460154
- [38] A. Da browski, Adsorption — from theory to practice, *Advanced Colloid and Interface Science*, Vol. 93, pp. 135-144, 2001.  
http://doi:10.1016/S0001-8686(00)00082-8
- [39] S.A.Umoren, I.B.Obot, E.E. Ebeno,N.O. Obi-Egbedib,The Inhibition of aluminium corrosion in hydrochloric acid solution by exudate gum from Raphiahookeri, *Desalination*,Vol. 247, pp. 561-572, 2009.  
http://doi:10.1016/j.desal.2008.09.005
- [40] Aboura, A., Seddak, A., Abascal, J., Effect of Microstructure on Crack Propagation of AISI304L Stainless Steel Pre-charged in Hydrogen, (2015) *International Review of Civil Engineering (IRECE)*, 6 (5), pp. 112-116.
- [41] Santos, P., Krambeck, L., Santos, D., Alves, T., Analysis of a Stainless Steel Heat Pipe based on Operation Limits, (2014) *International Review of Mechanical Engineering (IREME)*, 8 (3), pp. 599-608.
- [42] Kademani, A., Auti, A., Thermal and Structural Analysis of Stainless Steel 304 for Thermal Fatigue Testing Using ANSYS, (2014) *International Review of Mechanical Engineering (IREME)*, 8 (6), pp. 1116-1121.
- [43] Adam, S., Fairuz, M., Hussin, M., Hafiezal, M., Khaironisa, S., Investigate the Effect of Using Sunflower Oil as a Lubricant During Turning Operation of Stainless Steel, (2014) *International Review of Mechanical Engineering (IREME)*, 8 (1), pp. 22-27.

## Authors' information

Mechanical Engineering Department,  
Brawijaya University,  
Indonesia.



**Femiana Gapsari** was born in Bima, Indonesia, on 4th of July 1982. She obtained her bachelor's degree in material and metallurgy engineering Institut Teknologi Sepuluh Nopember Subaya Indonesia. She is currently a Ph.D. student under supervision of Prof. Rudy Soenoko. Her research is centered on Green Corrosion Inhibitors. In 2008 she received an offer to join Mechanical Engineering Department, Brawijaya University as lecturer. She has published her works in reputable journals indexed by scopus.



**Rudy Soenoko** is a professor of mechanical engineering , Brawijaya University. He was born on 11 of September 1949. He received the B.E. degree in mechanical engineering from the University of Brawijaya, Malang ,Indonesia and the M.Eng.Sc degrees in mechanical engineering from Melbourne University, Australia. Prof. Dr. Ir. Rudy Senoko , M.Eng. Sc. has more than 20 years experience in mechanical engineering, and he is the author of over 20 peer-reviewed scientific publications, conference papers, books and patents.



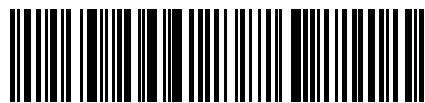
**Agus Suprapto** is a professor in failure analysis. He was born in Surabaya, Indonesia, on 7 of September 1958. He received the B.E. degree in mechanical engineering from the University of, Malang , Indonesia. The M.Sc and Ph.D. degree in mechanical engineering from Institut Teknologi Bandung, Indonesia and Universiti Teknologi Malaysia, respectively. He has published over 35 publications including, materials, energy conversion, construction, and manufacturers. He is Head of Institute of Research and Community Service Merdeka University, Indonesia.



**Wahyono Suprapto** is an associate professor in mechanical engineering department, Brawijaya University. He was born in Cilacap, Indonesia, on 17 of November 1955. He received the Ph.D. degree in Materials and Metallurgy engineering, University of Indonesia in 2011. Since January 1986, he has been with the Department of Mechanical Engineering, Brawijaya University. He worked research in powder metallurgy and mechanical testing. He has published over 10 publications and 2 books.



Praise Worthy Prize



1970-8742(201611)10:7;1-9